

## Thermodynamics Chemistry Study Guide Answers

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Thermodynamics is the study of the energy, principally heat energy, that accompanies chemical or physical changes. Some

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chemical reactions release heat energy; they are called exothermic reactions, and they have a negative enthalpy change. Others absorb heat energy and are called endothermic reactions, and they have a positive enthalpy change.

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## Thermodynamics Worksheet

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Thermodynamics Hugo Meekes, Solid State Chemistry, HG03.625, tel. 53200, H.Meekes@science.ru.nl Literature Book: Physical Chemistry; P.W. Atkins; edition 11, 10, 9 or 8, Oxford University Press Study guide: will be made available online during the course

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a state function that accounts for heat flow in chemical or physical processes that occurs at constant pressure when no work is performed except for PV work;  $H = E + PV$ ,  $\Delta H = \Delta E + P\Delta V = \text{heat at constant pressure (} q_{\text{subp}} \text{)} \rightarrow$  the change in enthalpy is equal to the energy flow as heat at constant P.

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released per mole =  $64.98\text{kJ} / 0.02125\text{mol} = 3,058\text{ kJ/mol}$  (one point) or,  $H_{\text{comb}} = -3,058\text{ kJ/mol}$ . Units not necessary (b)  $H_{\text{comb}} = -3,058\text{ kJ/mol} - 3,058\text{ kJ} = [6(-395.5) + 3(-285.85)] - [H^{\circ}\text{f phenol}]$  (one point)

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The experimental discovery that almost all chemical reactions either absorb or release heat led to the idea that all substances contain heat. Consequently, the heat of a reaction is the difference in the heat contents of the products and reactants:  $\Delta H = H_{\text{products}} - H_{\text{reactants}}$

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Book: Physical Chemistry; P.W. Atkins.; edition 11, 10, 9 or 8, Oxford University Press Study guide: This study guide. Aim

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Primary: Insight and skills Secondary: Knowledge Organisation  
Introduction Thermodynamics is built on two phenomenological perceptions, the first law and the second law describing the conditions for spontaneous processes.

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